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# 1-{(1*Z*)-1-[3-(2,4-Dichlorophenoxy)propoxy]-1-(2,4-difluorophenyl)prop-1-en-2-yl}-1*H*-1,2,4-triazole

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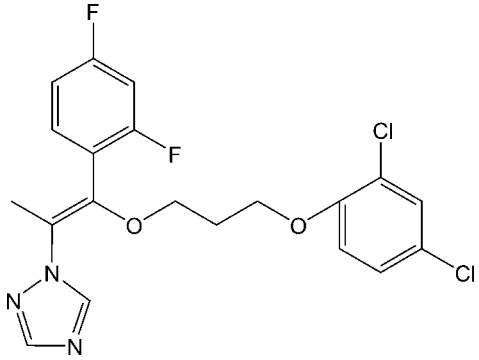
Key indicators: single-crystal X-ray study; $T$ = 293 K; mean $\sigma$(C–C) = 0.004 Å; $R$ factor = 0.054; $wR$ factor = 0.158; data-to-parameter ratio = 13.8.

In the title compound, $C_{20}H_{17}Cl_2F_2N_3O_2$, the triazole ring makes dihedral angles of 28.0 (3) and 72.5 (2)° with the 2,4-dichlorophenyl and 2,4-difluorophenyl rings, respectively, and the molecule adopts a *Z*-conformation about the C=C double bond. In the crystal, C—H···O and C—H···N hydrogen bonds link the molecules.

## Related literature

For a related structure and background to triazoles and further synthetic details, see: Shen *et al.* (2012).




## Experimental

### Crystal data

$C_{20}H_{17}Cl_2F_2N_3O_2$
$M_r$ = 440.27
Triclinic, $P\overline{1}$
$a$ = 7.4380 (15) Å
$b$ = 8.7600 (18) Å
$c$ = 15.892 (3) Å
$\alpha$ = 89.48 (3)°
$\beta$ = 84.57 (3)°
$\gamma$ = 73.74 (3)°
$V$ = 989.4 (3) Å$^3$
$Z$ = 2
Mo $K\alpha$ radiation
$\mu$ = 0.37 mm$^{-1}$
$T$ = 293 K
0.30 × 0.20 × 0.10 mm

### Data collection

Enraf–Nonius CAD-4 diffractometer
Absorption correction: $\psi$ scan (North *et al.*, 1968)
$T_{min}$ = 0.898, $T_{max}$ = 0.964
3933 measured reflections
3628 independent reflections
2564 reflections with $I > 2\sigma(I)$
$R_{int}$ = 0.022
3 standard reflections every 200 reflections
intensity decay: 1%

### Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.054
$wR(F^2)$ = 0.158
$S$ = 1.01
3628 reflections
262 parameters
2 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 1.21 e Å$^{-3}$
$\Delta\rho_{min}$ = −0.32 e Å$^{-3}$

**Table 1**
Hydrogen-bond geometry (Å, °).

| $D$—H···$A$ | $D$—H | H···$A$ | $D$···$A$ | $D$—H···$A$ |
|---|---|---|---|---|
| C4—H4$A$···O1$^{i}$ | 0.93 | 2.52 | 3.421 (4) | 163 |
| C17—H17$A$···N2$^{ii}$ | 0.93 | 2.56 | 3.400 (4) | 151 |

Symmetry codes: (i) $x-1, y, z$; (ii) $-x+1, -y+2, -z+1$.

Data collection: *CAD-4 EXPRESS* (Enraf–Nonius, 1994); cell refinement: *CAD-4 EXPRESS*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *PLATON* (Spek, 2009).

This research work was also supported by the Specialized Research Fund for the Doctoral Program of Higher Education (20113221110005).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB6698).

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Shen, F., Guo, S., Luan, Y.-Y., Wang, K. & Hu, Y.-H. (2012). *Acta Cryst.* E**68**, submited [HB6699].
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# supplementary materials

*Acta Cryst.* (2012). E**68**, o1780 [doi:10.1107/S1600536812017874]

## 1-{(1*Z*)-1-[3-(2,4-Dichlorophenoxy)propoxy]-1-(2,4-difluorophenyl)prop-1-en-2-yl}-1*H*-1,2,4-triazole

**Yuan-yuan Luan, Yong-hong Hu, Song Guo, Jing Zhu and Wen-ge Yang**

### Comment

As part of our studies on the synthesis of new triazole derivatives (Shen *et al.* 2012), the crystal structure of the title compound was determined.

In the molecular structure of the title compound the double bond is Z configurated. In the crystal, C-H···O and C-H···N hydrogen bonds link the molecules, in which they seem to be effective in the stabilization of the structure. (Table 1 and Fig. 2).

### Experimental

3 g (0.01 mol) 1-(2,4-difluorophenyl)-2-(1,2,4-triazol)-1-y1)propan-1-one, 10 g of a 50% aqueous sodium hydroxide, 15 ml toluene and 1.5 ml of a 40% aqueous solution of tetrabutyl ammonium hydroxide are mixed and heated to 323.15 K under vigorous stirring. 2.8g (0.01 mol) 1-bromo-3-(2.4-dichlorophenoxy)-propane, dissolved in 10 ml toluene, is instilled into the stirred and warmed solution in the course of 10 h. The mixture is subsequently stirred for another 20 h at 323.15 K. The reaction mixture is mixed with as much water and chloroform so that the aqueous phase becomes lighter than the organic phase. Thereafter, the organic and aqueous phases are separated. The organic phase is dried with sodium sulfate. The solvents are distilled under reduced pressure. The impure product herein is subsequently crystallized from a 1:1 mixture of ethyl acetate and ethanol. The purified product may be analytically identified as an approximately pure *Z*-isomer. Colourless blocks of the title compound were obtained by slow evaporation of an ethanol solution.

### Refinement

H atoms were positioned geometrically with C—H = 0.93 and 0.97 Å for aromatic and methylene H atoms, respectively, constrained to ride on their parent atoms, with $U_{iso}$(H) = 1.2 (or 1.5 for methyl groups) times $U_{eq}$(C).

### Computing details

Data collection: *CAD-4 EXPRESS* (Enraf–Nonius, 1994); cell refinement: *CAD-4 EXPRESS* (Enraf–Nonius, 1994); data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *PLATON* (Spek, 2009).

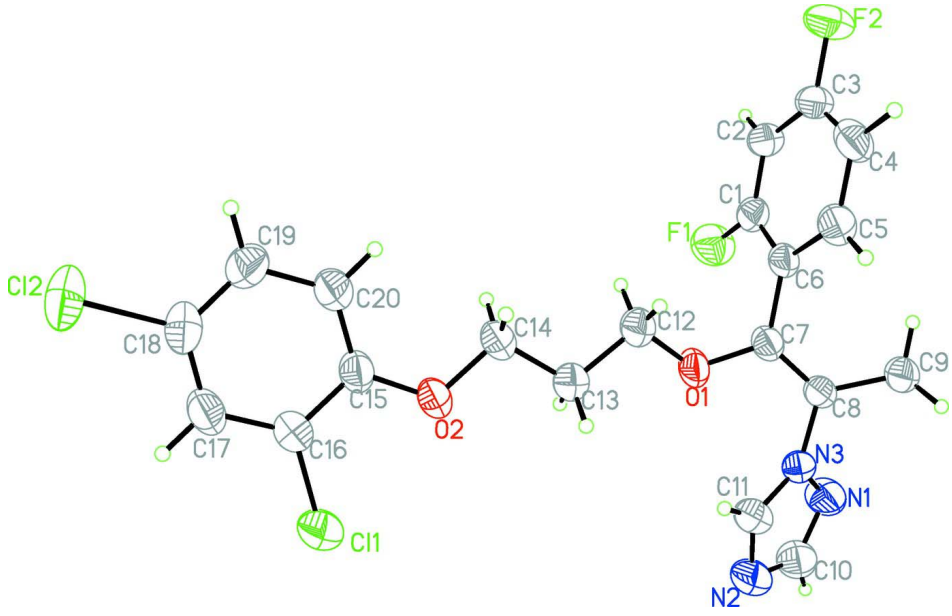


**Figure 1**

The molecular structure of the title molecule, with displacement ellipsoids drawn at 30% probability levels.

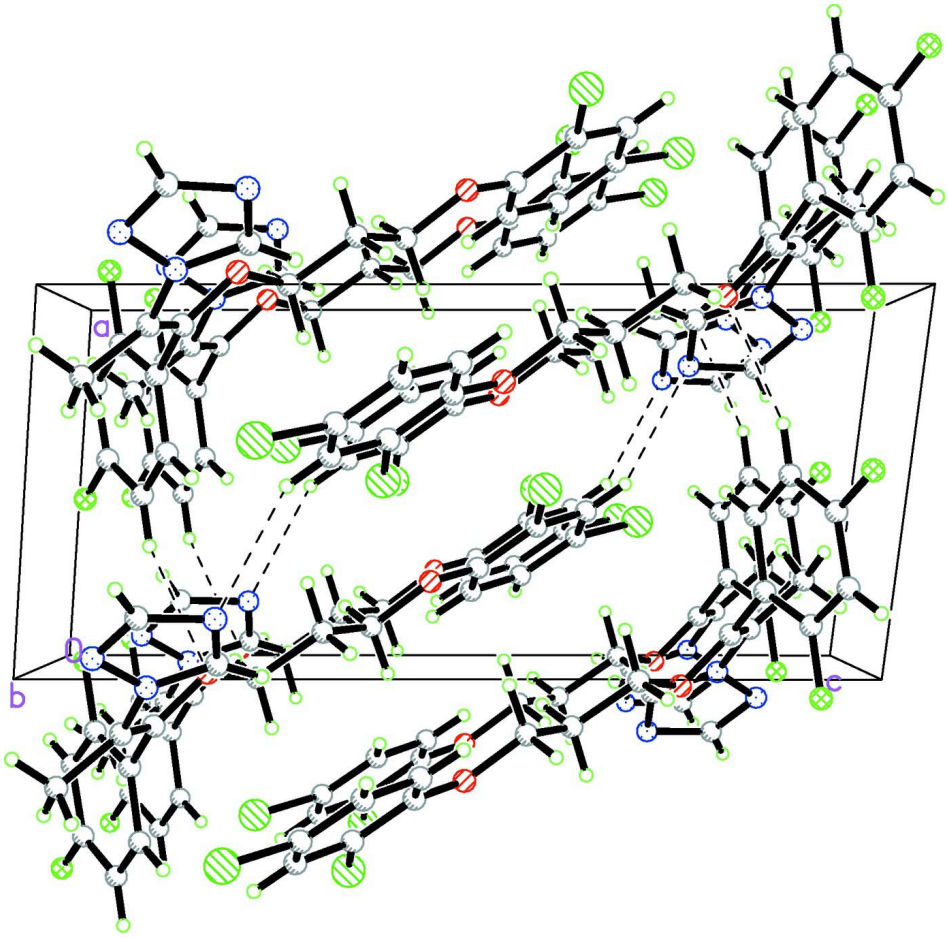


**Figure 2**

The packing diagram of the title compound. Hydron bonds are shown as dashed lines.

**1-{(1*Z*)-1-[3-(2,4-Dichlorophenoxy)propoxy]-1-(2,4-difluorophenyl)prop- 1-en-2-yl}-1*H*-1,2,4-triazole**

*Crystal data*

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$M_r$ = 440.27
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$a$ = 7.4380 (15) Å
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$\alpha$ = 89.48 (3)°
$\beta$ = 84.57 (3)°
$\gamma$ = 73.74 (3)°
$V$ = 989.4 (3) Å$^3$
$Z$ = 2
$F$(000) = 452
$D_x$ = 1.478 Mg m$^{-3}$
Mo $K\alpha$ radiation, $\lambda$ = 0.71073 Å
Cell parameters from 25 reflections
$\theta$ = 9–13°
$\mu$ = 0.37 mm$^{-1}$
$T$ = 293 K
Block, colorness
0.30 × 0.20 × 0.10 mm

*Data collection*

Enraf–Nonius CAD-4 diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
$\omega/2\theta$ scans
Absorption correction: $\psi$ scan ′(North *et al.*, 1968)
$T_{min}$ = 0.898, $T_{max}$ = 0.964
3933 measured reflections
3628 independent reflections

2564 reflections with $I > 2\sigma(I)$
$R_{int} = 0.022$
$\theta_{max} = 25.4°$, $\theta_{min} = 1.3°$
$h = 0 \rightarrow 8$
$k = -10 \rightarrow 10$
$l = -19 \rightarrow 19$
3 standard reflections every 200 reflections
intensity decay: 1%

*Refinement*

Refinement on $F^2$
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.054$
$wR(F^2) = 0.158$
$S = 1.01$
3628 reflections
262 parameters
2 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.1P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max} = 1.21$ e Å$^{-3}$
$\Delta\rho_{min} = -0.32$ e Å$^{-3}$

*Special details*

**Geometry**. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement**. Refinement of $F^2$ against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on $F^2$, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative $F^2$. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on $F^2$ are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)*

| | *x* | *y* | *z* | $U_{iso}$*/$U_{eq}$ |
|---|---|---|---|---|
| Cl1 | 0.48843 (15) | 0.85730 (11) | 0.59207 (6) | 0.0793 (3) |
| O1 | 0.0231 (3) | 1.0877 (2) | 0.22722 (12) | 0.0563 (6) |
| F1 | 0.0405 (2) | 0.8859 (2) | 0.07385 (13) | 0.0733 (5) |
| N1 | 0.0846 (4) | 1.4711 (3) | 0.11743 (16) | 0.0601 (7) |
| C1 | −0.1446 (4) | 0.9223 (3) | 0.09377 (16) | 0.0430 (6) |
| Cl2 | 0.39808 (13) | 0.30882 (11) | 0.71824 (5) | 0.0769 (3) |
| F2 | −0.5362 (3) | 0.7795 (2) | 0.05601 (13) | 0.0762 (6) |
| O2 | 0.2476 (3) | 0.8160 (2) | 0.46936 (12) | 0.0567 (5) |
| C2 | −0.2436 (4) | 0.8278 (3) | 0.06304 (18) | 0.0494 (7) |
| H2A | −0.1844 | 0.7386 | 0.0291 | 0.059* |
| N2 | 0.1804 (4) | 1.4856 (3) | 0.24639 (18) | 0.0661 (7) |
| N3 | 0.0067 (3) | 1.3826 (2) | 0.17311 (13) | 0.0412 (5) |
| C3 | −0.4343 (4) | 0.8715 (3) | 0.08482 (18) | 0.0502 (7) |
| C4 | −0.5252 (4) | 1.0021 (4) | 0.13428 (18) | 0.0536 (7) |
| H4A | −0.6547 | 1.0283 | 0.1479 | 0.064* |
| C5 | −0.4209 (4) | 1.0939 (3) | 0.16344 (18) | 0.0491 (7) |
| H5A | −0.4813 | 1.1835 | 0.1969 | 0.059* |
| C6 | −0.2260 (3) | 1.0556 (3) | 0.14398 (15) | 0.0391 (6) |
| C7 | −0.1090 (4) | 1.1539 (3) | 0.17330 (15) | 0.0403 (6) |
| C8 | −0.1167 (3) | 1.2983 (3) | 0.14444 (15) | 0.0388 (6) |
| C9 | −0.2440 (4) | 1.3857 (3) | 0.08190 (18) | 0.0517 (7) |
| H9A | −0.3213 | 1.3218 | 0.0657 | 0.078* |

| | | | | |
|---|---|---|---|---|
| H9B | −0.1700 | 1.4075 | 0.0329 | 0.078* |
| H9C | −0.3226 | 1.4840 | 0.1069 | 0.078* |
| C10 | 0.1855 (4) | 1.5294 (4) | 0.1652 (2) | 0.0635 (8) |
| H10A | 0.2557 | 1.5965 | 0.1441 | 0.076* |
| C11 | 0.0658 (4) | 1.3936 (3) | 0.24882 (19) | 0.0541 (7) |
| H11A | 0.0312 | 1.3434 | 0.2970 | 0.065* |
| C12 | 0.0011 (4) | 0.9622 (3) | 0.28022 (18) | 0.0532 (7) |
| H12A | 0.0406 | 0.8621 | 0.2488 | 0.064* |
| H12B | −0.1297 | 0.9811 | 0.3020 | 0.064* |
| C13 | 0.1210 (4) | 0.9566 (3) | 0.35150 (17) | 0.0516 (7) |
| H13A | 0.2475 | 0.9540 | 0.3288 | 0.062* |
| H13B | 0.0706 | 1.0520 | 0.3865 | 0.062* |
| C14 | 0.1277 (4) | 0.8136 (4) | 0.40447 (18) | 0.0542 (7) |
| H14A | 0.0024 | 0.8167 | 0.4293 | 0.065* |
| H14B | 0.1770 | 0.7173 | 0.3701 | 0.065* |
| C15 | 0.2795 (4) | 0.6921 (3) | 0.52364 (16) | 0.0463 (6) |
| C16 | 0.3907 (4) | 0.7002 (3) | 0.58833 (16) | 0.0483 (7) |
| C17 | 0.4295 (4) | 0.5822 (3) | 0.64753 (16) | 0.0536 (7) |
| H17A | 0.5028 | 0.5890 | 0.6909 | 0.064* |
| C18 | 0.3582 (4) | 0.4548 (4) | 0.64139 (17) | 0.0528 (7) |
| C19 | 0.2511 (4) | 0.4420 (4) | 0.57782 (19) | 0.0593 (8) |
| H19A | 0.2043 | 0.3545 | 0.5744 | 0.071* |
| C20 | 0.2131 (4) | 0.5605 (4) | 0.51874 (19) | 0.0563 (8) |
| H20A | 0.1418 | 0.5514 | 0.4750 | 0.068* |

*Atomic displacement parameters (Å²)*

| | $U^{11}$ | $U^{22}$ | $U^{33}$ | $U^{12}$ | $U^{13}$ | $U^{23}$ |
|---|---|---|---|---|---|---|
| Cl1 | 0.1089 (7) | 0.0746 (6) | 0.0782 (6) | −0.0529 (5) | −0.0461 (5) | 0.0141 (4) |
| O1 | 0.0545 (12) | 0.0545 (11) | 0.0728 (13) | −0.0284 (10) | −0.0339 (10) | 0.0327 (10) |
| F1 | 0.0522 (10) | 0.0749 (12) | 0.0908 (13) | −0.0154 (9) | −0.0029 (9) | −0.0087 (10) |
| N1 | 0.0683 (17) | 0.0634 (16) | 0.0633 (16) | −0.0394 (14) | −0.0173 (13) | 0.0194 (12) |
| C1 | 0.0434 (12) | 0.0415 (14) | 0.0459 (14) | −0.0140 (11) | −0.0079 (11) | 0.0103 (11) |
| Cl2 | 0.0869 (6) | 0.0753 (6) | 0.0652 (5) | −0.0163 (5) | −0.0137 (4) | 0.0343 (4) |
| F2 | 0.0778 (13) | 0.0781 (12) | 0.0946 (14) | −0.0494 (11) | −0.0350 (11) | 0.0096 (10) |
| O2 | 0.0700 (13) | 0.0610 (12) | 0.0528 (11) | −0.0325 (11) | −0.0326 (10) | 0.0198 (9) |
| C2 | 0.0566 (18) | 0.0445 (15) | 0.0531 (16) | −0.0213 (13) | −0.0140 (13) | 0.0054 (12) |
| N2 | 0.0677 (17) | 0.0615 (16) | 0.0811 (19) | −0.0301 (14) | −0.0314 (14) | 0.0039 (14) |
| N3 | 0.0425 (12) | 0.0370 (11) | 0.0479 (12) | −0.0154 (9) | −0.0102 (10) | 0.0041 (9) |
| C3 | 0.0591 (18) | 0.0484 (15) | 0.0554 (16) | −0.0294 (14) | −0.0247 (14) | 0.0148 (13) |
| C4 | 0.0390 (15) | 0.0676 (19) | 0.0604 (17) | −0.0222 (14) | −0.0149 (13) | 0.0189 (15) |
| C5 | 0.0455 (16) | 0.0507 (15) | 0.0533 (16) | −0.0151 (13) | −0.0114 (12) | 0.0080 (13) |
| C6 | 0.0379 (14) | 0.0400 (14) | 0.0435 (13) | −0.0149 (11) | −0.0134 (11) | 0.0133 (11) |
| C7 | 0.0399 (14) | 0.0448 (14) | 0.0393 (13) | −0.0144 (11) | −0.0117 (11) | 0.0072 (11) |
| C8 | 0.0400 (14) | 0.0402 (13) | 0.0404 (13) | −0.0158 (11) | −0.0115 (11) | 0.0051 (11) |
| C9 | 0.0614 (18) | 0.0445 (15) | 0.0576 (16) | −0.0222 (13) | −0.0270 (14) | 0.0173 (13) |
| C10 | 0.064 (2) | 0.0578 (18) | 0.081 (2) | −0.0335 (16) | −0.0185 (17) | 0.0128 (16) |
| C11 | 0.0612 (19) | 0.0527 (16) | 0.0555 (16) | −0.0227 (14) | −0.0211 (14) | 0.0025 (13) |
| C12 | 0.0614 (18) | 0.0523 (16) | 0.0539 (16) | −0.0245 (14) | −0.0218 (14) | 0.0183 (13) |
| C13 | 0.0611 (18) | 0.0471 (15) | 0.0523 (16) | −0.0202 (14) | −0.0196 (14) | 0.0103 (13) |

| | | | | | | |
|---|---|---|---|---|---|---|
| C14 | 0.0608 (18) | 0.0605 (17) | 0.0523 (16) | −0.0283 (15) | −0.0276 (14) | 0.0165 (14) |
| C15 | 0.0486 (16) | 0.0548 (16) | 0.0406 (14) | −0.0205 (13) | −0.0126 (12) | 0.0134 (12) |
| C16 | 0.0521 (17) | 0.0511 (16) | 0.0442 (15) | −0.0157 (13) | −0.0128 (12) | 0.0018 (12) |
| C17 | 0.0573 (18) | 0.0642 (18) | 0.0418 (15) | −0.0174 (15) | −0.0177 (13) | 0.0059 (13) |
| C18 | 0.0496 (16) | 0.0602 (18) | 0.0456 (15) | −0.0105 (14) | −0.0060 (13) | 0.0161 (13) |
| C19 | 0.0620 (19) | 0.0601 (18) | 0.0656 (19) | −0.0311 (15) | −0.0139 (15) | 0.0197 (15) |
| C20 | 0.0600 (19) | 0.0647 (18) | 0.0551 (17) | −0.0302 (15) | −0.0230 (14) | 0.0168 (14) |

*Geometric parameters (Å, º)*

| | | | |
|---|---|---|---|
| Cl1—C16 | 1.730 (3) | C7—C8 | 1.329 (3) |
| O1—C7 | 1.367 (3) | C8—C9 | 1.493 (3) |
| O1—C12 | 1.416 (3) | C9—H9A | 0.9600 |
| F1—C1 | 1.331 (3) | C9—H9B | 0.9600 |
| N1—C10 | 1.315 (4) | C9—H9C | 0.9600 |
| N1—N3 | 1.365 (3) | C10—H10A | 0.9300 |
| C1—C2 | 1.372 (4) | C11—H11A | 0.9300 |
| C1—C6 | 1.377 (4) | C12—C13 | 1.499 (4) |
| Cl2—C18 | 1.744 (3) | C12—H12A | 0.9700 |
| F2—C3 | 1.359 (3) | C12—H12B | 0.9700 |
| O2—C15 | 1.362 (3) | C13—C14 | 1.495 (4) |
| O2—C14 | 1.430 (3) | C13—H13A | 0.9700 |
| C2—C3 | 1.372 (4) | C13—H13B | 0.9700 |
| C2—H2A | 0.9300 | C14—H14A | 0.9700 |
| N2—C11 | 1.325 (4) | C14—H14B | 0.9700 |
| N2—C10 | 1.343 (4) | C15—C20 | 1.381 (4) |
| N3—C11 | 1.333 (3) | C15—C16 | 1.394 (4) |
| N3—C8 | 1.435 (3) | C16—C17 | 1.380 (4) |
| C3—C4 | 1.366 (4) | C17—C18 | 1.371 (4) |
| C4—C5 | 1.373 (4) | C17—H17A | 0.9300 |
| C4—H4A | 0.9300 | C18—C19 | 1.368 (4) |
| C5—C6 | 1.398 (4) | C19—C20 | 1.380 (4) |
| C5—H5A | 0.9300 | C19—H19A | 0.9300 |
| C6—C7 | 1.490 (3) | C20—H20A | 0.9300 |
| | | | |
| C7—O1—C12 | 120.2 (2) | N2—C11—N3 | 111.0 (3) |
| C10—N1—N3 | 102.4 (2) | N2—C11—H11A | 124.5 |
| F1—C1—C2 | 119.0 (3) | N3—C11—H11A | 124.5 |
| F1—C1—C6 | 117.2 (2) | O1—C12—C13 | 107.3 (2) |
| C2—C1—C6 | 123.8 (3) | O1—C12—H12A | 110.3 |
| C15—O2—C14 | 117.3 (2) | C13—C12—H12A | 110.3 |
| C1—C2—C3 | 116.7 (3) | O1—C12—H12B | 110.3 |
| C1—C2—H2A | 121.7 | C13—C12—H12B | 110.3 |
| C3—C2—H2A | 121.7 | H12A—C12—H12B | 108.5 |
| C11—N2—C10 | 102.3 (3) | C14—C13—C12 | 111.2 (2) |
| C11—N3—N1 | 108.8 (2) | C14—C13—H13A | 109.4 |
| C11—N3—C8 | 131.6 (2) | C12—C13—H13A | 109.4 |
| N1—N3—C8 | 119.6 (2) | C14—C13—H13B | 109.4 |
| F2—C3—C4 | 118.9 (3) | C12—C13—H13B | 109.4 |
| F2—C3—C2 | 118.1 (3) | H13A—C13—H13B | 108.0 |

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| C4—C3—C2 | 123.0 (3) | O2—C14—C13 | 107.4 (2) |
| C3—C4—C5 | 118.5 (3) | O2—C14—H14A | 110.2 |
| C3—C4—H4A | 120.8 | C13—C14—H14A | 110.2 |
| C5—C4—H4A | 120.8 | O2—C14—H14B | 110.2 |
| C4—C5—C6 | 121.5 (3) | C13—C14—H14B | 110.2 |
| C4—C5—H5A | 119.3 | H14A—C14—H14B | 108.5 |
| C6—C5—H5A | 119.3 | O2—C15—C20 | 125.4 (2) |
| C1—C6—C5 | 116.6 (2) | O2—C15—C16 | 116.3 (2) |
| C1—C6—C7 | 120.4 (2) | C20—C15—C16 | 118.3 (2) |
| C5—C6—C7 | 122.9 (2) | C17—C16—C15 | 121.0 (3) |
| C8—C7—O1 | 118.3 (2) | C17—C16—Cl1 | 119.6 (2) |
| C8—C7—C6 | 123.1 (2) | C15—C16—Cl1 | 119.3 (2) |
| O1—C7—C6 | 118.4 (2) | C18—C17—C16 | 118.9 (3) |
| C7—C8—N3 | 119.7 (2) | C18—C17—H17A | 120.5 |
| C7—C8—C9 | 125.8 (2) | C16—C17—H17A | 120.5 |
| N3—C8—C9 | 114.4 (2) | C19—C18—C17 | 121.5 (3) |
| C8—C9—H9A | 109.5 | C19—C18—Cl2 | 119.1 (2) |
| C8—C9—H9B | 109.5 | C17—C18—Cl2 | 119.4 (2) |
| H9A—C9—H9B | 109.5 | C18—C19—C20 | 119.3 (3) |
| C8—C9—H9C | 109.5 | C18—C19—H19A | 120.4 |
| H9A—C9—H9C | 109.5 | C20—C19—H19A | 120.4 |
| H9B—C9—H9C | 109.5 | C19—C20—C15 | 121.0 (3) |
| N1—C10—N2 | 115.5 (3) | C19—C20—H20A | 119.5 |
| N1—C10—H10A | 122.2 | C15—C20—H20A | 119.5 |
| N2—C10—H10A | 122.2 | | |

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| F1—C1—C2—C3 | −179.4 (2) | C11—N3—C8—C9 | 141.9 (3) |
| C6—C1—C2—C3 | −0.1 (4) | N1—N3—C8—C9 | −38.1 (3) |
| C10—N1—N3—C11 | −0.2 (3) | N3—N1—C10—N2 | 0.4 (4) |
| C10—N1—N3—C8 | 179.8 (2) | C11—N2—C10—N1 | −0.4 (4) |
| C1—C2—C3—F2 | −179.4 (2) | C10—N2—C11—N3 | 0.3 (3) |
| C1—C2—C3—C4 | 0.2 (4) | N1—N3—C11—N2 | −0.1 (3) |
| F2—C3—C4—C5 | 179.7 (2) | C8—N3—C11—N2 | 179.9 (2) |
| C2—C3—C4—C5 | 0.0 (4) | C7—O1—C12—C13 | −159.4 (2) |
| C3—C4—C5—C6 | −0.4 (4) | O1—C12—C13—C14 | −172.2 (2) |
| F1—C1—C6—C5 | 179.1 (2) | C15—O2—C14—C13 | −177.7 (2) |
| C2—C1—C6—C5 | −0.2 (4) | C12—C13—C14—O2 | 178.8 (3) |
| F1—C1—C6—C7 | 0.6 (3) | C14—O2—C15—C20 | 3.4 (4) |
| C2—C1—C6—C7 | −178.7 (2) | C14—O2—C15—C16 | −177.6 (3) |
| C4—C5—C6—C1 | 0.5 (4) | O2—C15—C16—C17 | 179.1 (2) |
| C4—C5—C6—C7 | 178.9 (2) | C20—C15—C16—C17 | −1.8 (4) |
| C12—O1—C7—C8 | 159.7 (3) | O2—C15—C16—Cl1 | −3.3 (4) |
| C12—O1—C7—C6 | −25.5 (4) | C20—C15—C16—Cl1 | 175.8 (2) |
| C1—C6—C7—C8 | 109.1 (3) | C15—C16—C17—C18 | 0.7 (4) |
| C5—C6—C7—C8 | −69.4 (4) | Cl1—C16—C17—C18 | −177.0 (2) |
| C1—C6—C7—O1 | −65.5 (3) | C16—C17—C18—C19 | 0.4 (5) |
| C5—C6—C7—O1 | 116.1 (3) | C16—C17—C18—Cl2 | −177.4 (2) |
| O1—C7—C8—N3 | −2.7 (4) | C17—C18—C19—C20 | −0.3 (5) |
| C6—C7—C8—N3 | −177.2 (2) | Cl2—C18—C19—C20 | 177.5 (2) |

| | | | |
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| O1—C7—C8—C9 | 176.6 (2) | C18—C19—C20—C15 | −0.9 (5) |
| C6—C7—C8—C9 | 2.0 (4) | O2—C15—C20—C19 | −179.0 (3) |
| C11—N3—C8—C7 | −38.8 (4) | C16—C15—C20—C19 | 1.9 (5) |
| N1—N3—C8—C7 | 141.2 (3) | | |

*Hydrogen-bond geometry (Å, º)*

| *D*—H···*A* | *D*—H | H···*A* | *D*···*A* | *D*—H···*A* |
|---|---|---|---|---|
| C4—H4*A*···O1$^{i}$ | 0.93 | 2.52 | 3.421 (4) | 163 |
| C17—H17*A*···N2$^{ii}$ | 0.93 | 2.56 | 3.400 (4) | 151 |

Symmetry codes: (i) *x*−1, *y*, *z*; (ii) −*x*+1, −*y*+2, −*z*+1.